ionic conductivity
Defined for ionic species B by

$$\lambda = |z_B| F u_B$$

where z_B is the charge number of the ionic species B, F is the Faraday constant, and u_B is the *electric mobility* of species B. In most current practice z_B is taken as unity, i.e. ionic conductivity is taken as that of species such as Na^+, $Ca^{2+}/2$, $La^{3+}/3$ etc. To avoid ambiguity the species considered should be clearly stated, e.g. as $\lambda(Ca^{2+}/2)$.

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