

**protonation constant**

The *equilibrium constant*  $K_{Hn}$  for the addition of the  $n$ th proton to a charged or uncharged ligand. The cumulative protonation constant,  $\beta_{Hn}$ , is the equilibrium constant for the formation of  $H_nL$  from  $nH^+$  and  $L$ .

The same equilibrium constant may be described in several ways, depending on how the ligand,  $L$ , or the acid, has been defined.

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